UNIVERSITY OF MYSORE

Postgraduate Entrance Examination August - 2024

QUESTION PAPER BOOKLET NO.

Entrance Reg. No.						

SUBJECT CODE : 11

QUESTION BOOKLET

(Read carefully the instructions given in the Question Booklet)

COURSE : M.Sc. **SUBJECT : GROUP-4 : CHEMISTRY/ORGANIC CHEMISTRY**

MAXIMUM MARKS : 50 **MAXIMUM TIME : 75 MINUTES**

(Including time for filling O.M.R. Answer sheet)

INSTRUCTIONS TO THE CANDIDATES

1. The sealed question paper booklet containing 50 questions enclosed with O.M.R. Answer Sheet is given to you.
2. Verify whether the given question booklet is of the same subject which you have opted for examination.
3. Open the question paper seal carefully and take out the enclosed O.M.R. Answer Sheet outside the question booklet and fill up the general information in the O.M.R. Answer sheet. If you fail to fill up the details in the form as instructed, you will be personally responsible for consequences arising during evaluating your Answer Sheet.
4. During the examination:
   a) Read each question carefully.
   b) Determine the Most appropriate/correct answer from the four available choices given under each question.
   c) Completely darken the relevant circle against the Question in the O.M.R. Answer Sheet. For example, in the question paper if "C" is correct answer for Question No.8, then darken against Sl. No.8 of O.M.R. Answer Sheet using Blue/Black Ball Point Pen as follows:

   Question No. 8. Ⓐ Ⓑ ● Ⓓ (Only example) (Use Ball Pen only)
5. Rough work should be done only on the blank space provided in the Question Booklet. Rough work should not be done on the O.M.R. Answer Sheet.
6. If more than one circle is darkened for a given question, such answer is treated as wrong and no mark will be given. See the example in the O.M.R. Sheet.
7. The candidate and the Room Supervisor should sign in the O.M.R. Sheet at the specified place.
8. Candidate should return the original O.M.R. Answer Sheet and the university copy to the Room Supervisor after the examination.
9. Candidate can carry the question booklet and the candidate copy of the O.M.R. Sheet.
10. The calculator, pager and mobile phone are not allowed inside the examination hall.
11. If a candidate is found committing malpractice, such a candidate shall not be considered for admission to the course and action against such candidate will be taken as per rules.
12. Candidates have to get qualified in the respective entrance examination by securing a minimum of 8 marks in case of SC/ST/Cat-I Candidates, 9 marks in case of OBC Candidates and 10 marks in case of other Candidates out of 50 marks.

INSTRUCTIONS TO FILL UP THE O.M.R. SHEET

1. There is only one most appropriate/correct answer for each question.
2. For each question, only one circle must be darkened with BLUE or BLACK ball point pen only. Do not try to alter it.
3. Circle should be darkened completely so that the alphabet inside it is not visible.
4. Do not make any unnecessary marks on O.M.R. Sheet.
5. Mention the number of questions answered in the appropriate space provided in the O.M.R. sheet otherwise O.M.R. sheet will not be subjected for evaluation.

ಗಮನಿಸಿ : ಸೂಚನೆಗಳ ಕನ್ನಡ ಆವೃತ್ತಿಯು ಈ ಪುಸ್ತಕದ ಹಿಂಭಾಗದಲ್ಲಿ ಮುದ್ರಿಸಲ್ಪಟ್ಟಿದೆ.

1. The hydrogen emission spectrum comprises radiation of discrete :

(A) Frequency (B) Wavelength

(C) Energy (D) None of the above

2. The character of p-orbital which determines the geometry of molecules is,

(A) Non directional (B) Directional

(C) Vertical (D) Horizontal

3. The shielding constant experienced by a 2*p* electron in the nitrogen atom is,

(A) 4.9 (B) 8.1

(C) 10.9 (D) 3.10

4. Consider the following statements:

I. In ionic bond formation, electropositive element donate electron to electronegative element

II. Ionic bonds are quite strong and they are omnidirectional.

(A) Both statement I and II are true

(B) Statement I is true and statement II is false

(C) Statement I is false and statement II is true

(D) Both statement I and II are false

5. The correct statement/s among the following for fluorite structure is/are

I. In fluorite structure Ca^{2+} ion is surrounded by six F^{-} ions.

II. The coordination number in fluorite structure is 6 and 4, so this called an 6:4 arrangement.

III. Ca^{2+} ions form hexagonal-close packed arrangement.

(A) I and II only (B) II and III only

(C) I, II and III (D) None of the above

6. Match the following for the shape of the molecules;

I		II	
a)	I_3^-	i)	Linear
b)	NH_4^+	ii)	Tetrahedral
c)	I_3^+	iii)	Bent
d)	SF_4	iv)	Distorted tetrahedral

(A) a-i, b-ii, c-iii, d-iv (B) a-i, b-iv, c-iii, d-ii

(C) a-iii, b-iv, c-ii, d-i (D) a-iii, b-ii, c-i, d-iv

7. The hybridization of $[Co(NH_3)_6]^{3+}$, $[CoCl_6]^{3-}$ and $[Ni(CN)_4]^{2-}$ respectively are,

(A) d^2sp^3, sp^3d^2, sp^3 (B) sp^3d^2, d^2sp^3, dsp^2

(C) d^2sp^3, sp^3d^2, dsp^2 (D) sp^3d^2, d^2sp^3, sp^3

8. Using MO theory, arrange the following species in the increasing order of their bond length : O_2, O_2^+, O_2^{2+} O_2^-, O_2^{2-}

(A) $O_2 < O_2^+ < O_2^{2+} < O_2^- < O_2^{2-}$ (B) $O_2^+ > O_2 > O_2^{2+} > O_2^- > O_2^{2-}$

(C) $O_2 < O_2^+ < O_2^{2+} < O_2^- < O_2^{2-}$ (D) $O_2^{2+} < O_2^+ < O_2 < O_2^- < O_2^{2-}$

9. Number of antibonding electrons in CO molecule is,

(A) 0 (B) 2

(C) 4 (D) 6

10. Which of the following is the correct order of their stability?

(A) $CN < NO^+ < NO^-$ (B) $NO^- < CN < NO^+$

(C) $NO^+ < CN < NO^-$ (D) $CN < NO^- < NO^+$

11. The IUPAC nomenclature of $Na_3[Ag(S_2O_3)_2]$ is,

(A) Sodium di(thiouslphato)silver(II)

(B) Sodium bis(thiosulhato)argenate(I)

(C) Sodium bis(thiosulphato)argenate(III)

(D) Bis(thiosulphato)argenate(I)sodium

12. The complexes $[Co(NH_3)_5NO_2]Cl_2$ and $[Co(NH_3)_5ONO]Cl_2$ are an example of

(A) Linkage isomers
(B) Ionisation isomers
(C) Geometrical isomers
(D) Coordination isomers

13. The complex with spin-only magnetic moment of 4.9 BM is

(A) $[Fe(H_2O)_6]^{2+}$
(B) $[Fe(CN)_6]^{3-}$
(C) $[Fe(CN)_6]^{4-}$
(D) $[Fe(H_2O)_6]^{3+}$

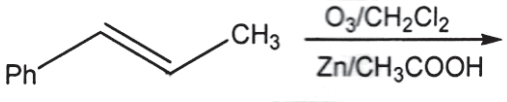
14. The CFSE for octahedral (Δ_o) and tetrahedral (Δ_t) complexes is related as

(A) $\Delta_t \simeq 4/9\ \Delta_o$
(B) $\Delta_t \simeq ½\ \Delta_o$
(C) $\Delta_o \simeq 2\ \Delta_t$
(D) $\Delta_o \simeq 4/9\ \Delta_t$

15. In $[FeF_6]^{3-}$ and $[CoF_6]^{3-}$

(A) $[FeF_6]^{3-}$ is coloured and $[CoF_6]^{3-}$ is colourless
(B) $[FeF_6]^{3-}$ is colourless and $[CoF6]^{3-}$ is coloured
(C) $[FeF_6]^{3-}$ and $[CoF_6]^{3-}$ are both colourless
(D) $[FeF_6]^{3-}$ and $[CoF_6]^{3-}$ are both coloured

16. Which one of the following is the major product of the reaction?



Major product

(A) PhCHO + CH_3CHO
(B) PhCHO + CH_3COOH
(C) PhCOOH + CH_3COOH
(D) PhCHO + HCHO

17. Which of the following statement is true for sucrose?

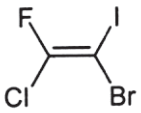
(A) Reducing sugar
(B) Non- reducing sugar
(C) Undergoes Mutarotation
(D) It is a monosaccharide

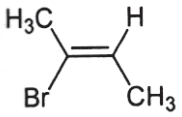
18. Sulphonation of thiophene gives

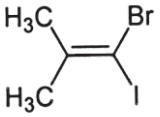
(A) Thiophene-3-sulphonic acid (B) Thiophene-4-sulphonic acid

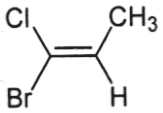
(C) Thiophene-2-sulphonic acid (D) Thiophene-2,4-disulphonic acid

19. Which of the following compound do not show geometrical isomerism?

(A) 

(B) 

(C) 

(D) 

20. When cyclohexanone oxime is treated with H_2SO_4 the ring expansion take place which is due to

(A) Curtius rearrangement (B) Beckmann rearrangement

(C) Favorski rearrangement (D) Claisen rearrangement

21. Which one of the following product is formed in Gatterman-Koch reaction?

(A) Benzoic acid (B) Benzophenone

(C) Benzoyl chloride (D) Benzaldehyde

22. Paal-Knor method is used to synthesise the following heterocyclic compound

(A) Pyridine (B) Quinoline

(C) Indole (D) Furan

23. Which of the statement is true for aromatic compounds?

(A) Cyclic (B) Planar

(C) Obeys Huckel rule (D) All of these

24. What kind of product(s) is formed in S_N1 reaction?

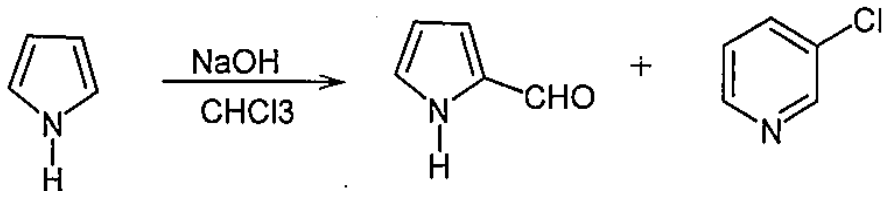
(A) Inversion product

(B) Retention of configuration product

(C) Racemic mixture product

(D) Isomerization product

25. The following conversion is an example of



(A) Arndt-Eistert homologation

(B) Mannich reaction

(C) Riemer Tiemann reaction addition

(D) Chichibabin reaction

26. Racemic modification can be resolved by

(A) The use of Enzymes

(B) Mechanical Method

(C) Chemical Method

(D) All of the above

27. Which of the following reducing agent is used for the conversion of keto (C = O) into the $-CH_2-$ group?

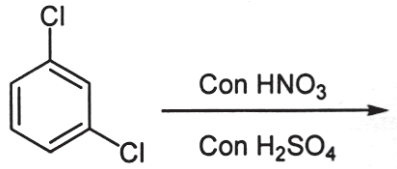
(A) Zn/Hg, HCl

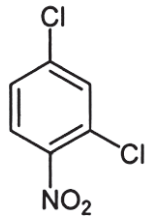
(B) $NaBH_4$

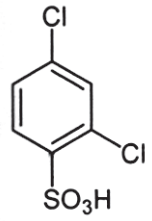
(C) $LiAlH_4$

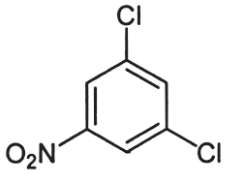
(D) 1, 3-BH_3; H_2O_2/NaOH

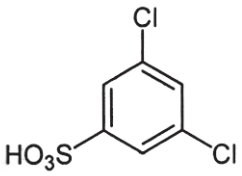
28. The major product formed in the following reaction is



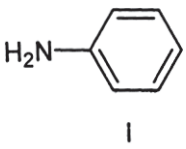
(A) 

(B) 

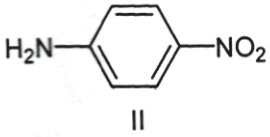
(C) 

(D) 

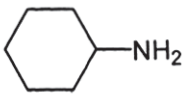
29. Rank the following compounds in the order of decreasing basicity:



I



II



III

(A) I, II, III
(B) III, II, I
(C) II, I, III
(D) III, I, II

30. The priority sequence for assigning R, S configuration of lactic acid is

(A) $-CH_3$, –COOH, –OH, –H
(B) –COOH, –OH, $-CH_3$, –H
(C) –OH, $-CH_3$, –COOH, –H
(D) –OH, –COOH, $-CH_3$, –H

31. Kinetic energy of a single molecule is given by the expression;

(A) RT
(B) RT/N
(C) nRT
(D) 1.5RT/N

32. Molar refraction is an,

(A) Additive property
(B) Constitutive property
(C) Additive and constitutive property
(D) Thermodynamic property

33. A container of volume 5.0 L is divided into two compartments of equal size. In the left compartment there is nitrogen at 1.0 atmosphere pressure and 25 °C, in the right compartment there is hydrogen at the same temperature and pressure. What will happen when the partition is removed?

(A) The entropy increases and the free energy decreases
(B) The entropy decreases and the free energy decreases
(C) The entropy increases and the free energy increases
(D) The entropy decreases and the free energy increases

34. Plants and living beings are the examples of,

(A) Isolated system
(B) Adiabatic system
(C) Open system
(D) Closed system

35. Which of the following is the primary reference electrode for the measurement of electrode potential?

(A) Glass electrode
(B) Normal calomel electrode
(C) Standard hydrogen electrode
(D) Silver-silver chloride electrode

36. In a second order reaction $2A \rightarrow$ products, if the concentration of A is doubled, $t_{1/2}$ of the reaction is,

(A) Doubled
(B) Quadrupled
(C) Halved
(D) Unchanged

37. In physical adsorption, the forces of attraction are,

(A) Ionic
(B) Covalent
(C) Vander Waal's
(D) H-bonding

38. The rate constant for the second-order decomposition of N_sO follows the following equation: $(k = 5.00 \times 10^{11}\ dm^3\ mol^{-1}\ s^{-1})\ exp\ (-29{,}000\ K/T)$. The activation energy of the reaction is,

(A) 5.76 kJ mol^{-1}
(B) 24.1 kJ mol^{-1}
(C) 57.6 kJ mol^{-1}
(D) 241.1 kJ mol^{-1}

39. The molecular weight of NaCl, assuming 100% dissociation in solution, as determined by elevation of boiling point method is,

(A) equal to 58.5 (B) 29.25

(C) greater than 58.5 (D) Zero

40. The phase diagram of KI-H_2O is of simple eutectic type. The eutectic composition is 52.0 weight% KI and it freezes at –22.0 °C. The phases present in a solution containing 25 weight% of KI at –18 °C are;

(A) Ice + KI solution (B) Ice + solid KI

(C) KI + pure water (D) KI + saturated KI solution

41. A concentration cell consisting of two zinc electrodes are immersed in 0.01 molal and 0.1 molal $ZnSO_4$ solutions at 25 °C. The two solutions are separated by a salt bridge. The EMF of the cell is (the activity coefficient of the electrolyte may be assumed to be unity)

(A) 0.0147 V (B) 0.0591 V

(C) 0.0295 V (D) 0.1182V

42. The conductivity of 0.01 M NaCl solution is 0.00147 ohm-1 cm-1. When 50 ml of water is added to the above solution, its conductivity

(A) Increases

(B) Decreases

(C) Remains unchanged

(D) First increases and then decreases

43. The EMF of Daniel cell at 298 K is E_1, $Zn|ZnSO_4$ (1 M) || $CuSO_4$ (1 M)|Cu, when the concentration of $ZnSO_4$ is 1.0 M and that of $CuSO_4$ is 0.01 M the EMF changed to E_2. The relationship between E_1 and E_2 is,

(A) $E_2 = 0 \neq E_1$ (B) $E_1 > E_2$

(C) $E_1 < E_2$ (D) $E_1 = E_2$

44. The hydrogen ion concentration of a solution with pH value 2.69 is given by,

(A) 2.29×10^{-3} M (B) 3.69×10^{-2} M

(C) 4.31×10^{-4} M (D) 2.042×10^{-3} M

45. The substances when exposed to light radiations of short wave length emit light of longer wave length are called as

(A) Photosensitized substance (B) Phosphorescent substance

(C) Fluorescent substance (D) Non radiative substance

46. In solvent extraction, which of the following term describes the ratio of the concentration of solute in the organic phase to its concentration in the aqueous phase?

(A) Molarity (B) Solubility product

(C) Partition coefficient (D) Stoichiometry

47. An error that consistently occurs in the same direction in repeated measurements is known as:

(A) Random error (B) Systematic error

(C) Absolute error (D) Relative error

48. The R_f value in thin layer chromatography (TLC) is calculated as:

(A) Distance travelled by sample to Distance travelled by solvent front

(B) Distance travelled by solvent front to Distance travelled by sample

(C) Distance travelled by stationary phase to Distance travelled by mobile phase

(D) None of the above

49. Which of the following is the application of redox titrations?

(A) Measurement of pH

(B) Detection of alkali metals

(C) Determination of hardness of water

(D) Quantification of organic compounds

50. Which type of titration curve exhibits a sharp equivalence point?

(A) Strong acid Vs Strong base (B) Weak acid Vs Strong base

(C) Weak base Vs weak acid (D) None of the above

▽▽▽▽

Rough Work

ಅಭ್ಯರ್ಥಿಗಳಿಗೆ ಸೂಚನೆಗಳು

1. ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯ ಜೊತೆಗೆ 50 ಪ್ರಶ್ನೆಗಳನ್ನು ಹೊಂದಿರುವ ಮೊಹರು ಮಾಡಿದ ಪ್ರಶ್ನೆ ಪುಸ್ತಕವನ್ನು ನಿಮಗೆ ನೀಡಲಾಗಿದೆ.
2. ಕೊಟ್ಟಿರುವ ಪ್ರಶ್ನೆ ಪುಸ್ತಕವು, ನೀವು ಪರೀಕ್ಷೆಗೆ ಆಯ್ಕೆ ಮಾಡಿಕೊಂಡಿರುವ ವಿಷಯಕ್ಕೆ ಸಂಬಂಧಿಸಿದ್ದೇ ಎಂಬುದನ್ನು ಪರಿಶೀಲಿಸಿರಿ.
3. ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯ ಮೊಹರನ್ನು ಜಾಗ್ರತೆಯಿಂದ ತೆರೆಯಿರಿ ಮತ್ತು ಪ್ರಶ್ನೆಪತ್ರಿಕೆಯಿಂದ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯನ್ನು ಹೊರಗೆ ತೆಗೆದು, ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯಲ್ಲಿ ಸಾಮಾನ್ಯ ಮಾಹಿತಿಯನ್ನು ತುಂಬಿರಿ. ಕೊಟ್ಟಿರುವ ಸೂಚನೆಯಂತೆ ನೀವು ನಮೂನೆಯಲ್ಲಿನ ವಿವರಗಳನ್ನು ತುಂಬಲು ವಿಫಲರಾದರೆ, ನಿಮ್ಮ ಉತ್ತರ ಹಾಳೆಯ ಮೌಲ್ಯಮಾಪನ ಸಮಯದಲ್ಲಿ ಉಂಟಾಗುವ ಪರಿಣಾಮಗಳಿಗೆ ವೈಯಕ್ತಿಕವಾಗಿ ನೀವೇ ಜವಾಬ್ದಾರರಾಗಿರುತ್ತೀರಿ.
4. ಪರೀಕ್ಷೆಯ ಸಮಯದಲ್ಲಿ:
   a) ಪ್ರತಿಯೊಂದು ಪ್ರಶ್ನೆಯನ್ನು ಜಾಗ್ರತೆಯಿಂದ ಓದಿರಿ.
   b) ಪ್ರತಿ ಪ್ರಶ್ನೆಯ ಕೆಳಗೆ ನೀಡಿರುವ ನಾಲ್ಕು ಲಭ್ಯ ಆಯ್ಕೆಗಳಲ್ಲಿ ಅತ್ಯಂತ ಸರಿಯಾದ/ ಸೂಕ್ತವಾದ ಉತ್ತರವನ್ನು ನಿರ್ಧರಿಸಿ.
   c) ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿನ ಸಂಬಂಧಿಸಿದ ಪ್ರಶ್ನೆಯ ವೃತ್ತಾಕಾರವನ್ನು ಸಂಪೂರ್ಣವಾಗಿ ತುಂಬಿರಿ. ಉದಾಹರಣೆಗೆ, ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯಲ್ಲಿ ಪ್ರಶ್ನೆ ಸಂಖ್ಯೆ 8ಕ್ಕೆ "C" ಸರಿಯಾದ ಉತ್ತರವಾಗಿದ್ದರೆ, ನೀಲಿ/ಕಪ್ಪು ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್ ಬಳಸಿ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯ ಕ್ರಮ ಸಂಖ್ಯೆ 8ರ ಮುಂದೆ ಈ ಕೆಳಗಿನಂತೆ ತುಂಬಿರಿ:

   ಪ್ರಶ್ನೆ ಸಂಖ್ಯೆ 8. Ⓐ Ⓑ ● Ⓓ (ಉದಾಹರಣೆ ಮಾತ್ರ) (ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್ ಮಾತ್ರ ಉಪಯೋಗಿಸಿ)
5. ಉತ್ತರದ ಪೂರ್ವಸಿದ್ಧತೆಯ ಬರವಣಿಗೆಯನ್ನು (ಚಿತ್ತು ಕೆಲಸ) ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯಲ್ಲಿ ಒದಗಿಸಿದ ಖಾಲಿ ಜಾಗದಲ್ಲಿ ಮಾತ್ರವೇ ಮಾಡಬೇಕು (ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯಲ್ಲಿ ಮಾಡಬಾರದು).
6. ಒಂದು ನಿರ್ದಿಷ್ಟ ಪ್ರಶ್ನೆಗೆ ಒಂದಕ್ಕಿಂತ ಹೆಚ್ಚು ವೃತ್ತಾಕಾರವನ್ನು ಗುರುತಿಸಲಾಗಿದ್ದರೆ, ಅಂತಹ ಉತ್ತರವನ್ನು ತಪ್ಪು ಎಂದು ಪರಿಗಣಿಸಲಾಗುತ್ತದೆ ಮತ್ತು ಯಾವುದೇ ಅಂಕವನ್ನು ನೀಡಲಾಗುವುದಿಲ್ಲ. ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿನ ಉದಾಹರಣೆ ನೋಡಿ.
7. ಅಭ್ಯರ್ಥಿ ಮತ್ತು ಕೊಠಡಿ ಮೇಲ್ವಿಚಾರಕರು ನಿರ್ದಿಷ್ಟಪಡಿಸಿದ ಸ್ಥಳದಲ್ಲಿ ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯ ಮೇಲೆ ಸಹಿ ಮಾಡಬೇಕು.
8. ಅಭ್ಯರ್ಥಿಯು ಪರೀಕ್ಷೆಯ ನಂತರ ಕೊಠಡಿ ಮೇಲ್ವಿಚಾರಕರಿಗೆ ಮೂಲ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆ ಮತ್ತು ವಿಶ್ವವಿದ್ಯಾನಿಲಯದ ಪ್ರತಿಯನ್ನು ಹಿಂದಿರುಗಿಸಬೇಕು.
9. ಅಭ್ಯರ್ಥಿಯು ಪ್ರಶ್ನೆ ಪುಸ್ತಕವನ್ನು ಮತ್ತು ಓ.ಎಂ.ಆರ್. ಅಭ್ಯರ್ಥಿಯ ಪ್ರತಿಯನ್ನು ತಮ್ಮ ಜೊತೆ ತೆಗೆದುಕೊಂಡು ಹೋಗಬಹುದು.
10. ಕ್ಯಾಲ್ಕುಲೇಟರ್, ಪೇಜರ್ ಮತ್ತು ಮೊಬೈಲ್ ಫೋನ್‌ಗಳನ್ನು ಪರೀಕ್ಷಾ ಕೊಠಡಿಯ ಒಳಗೆ ಅನುಮತಿಸಲಾಗುವುದಿಲ್ಲ.
11. ಅಭ್ಯರ್ಥಿಯು ದುಷ್ಕೃತ್ಯದಲ್ಲಿ ತೊಡಗಿರುವುದು ಕಂಡುಬಂದರೆ, ಅಂತಹ ಅಭ್ಯರ್ಥಿಯನ್ನು ಕೋರ್ಸ್‌ಗೆ ಪರಿಗಣಿಸಲಾಗುವುದಿಲ್ಲ ಮತ್ತು ನಿಯಮಗಳ ಪ್ರಕಾರ ಅಂತಹ ಅಭ್ಯರ್ಥಿಯ ವಿರುದ್ಧ ಕ್ರಮ ಕೈಗೊಳ್ಳಲಾಗುವುದು.
12. ಈ ಪ್ರವೇಶ ಪರೀಕ್ಷೆಯಲ್ಲಿ ಅರ್ಹರಾಗಲು ಒಟ್ಟು 50 ಅಂಕಗಳಲ್ಲಿ SC/ST/Cat-I ಅಭ್ಯರ್ಥಿಗಳು ಕನಿಷ್ಠ 8 ಅಂಕಗಳನ್ನು, OBC ಅಭ್ಯರ್ಥಿಗಳು ಕನಿಷ್ಠ 9 ಅಂಕಗಳನ್ನು ಮತ್ತು ಇನ್ನಿತರ ಅಭ್ಯರ್ಥಿಗಳು ಕನಿಷ್ಠ 10 ಅಂಕಗಳನ್ನು ಪಡೆಯತಕ್ಕದ್ದು.

ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯನ್ನು ತುಂಬಲು ಸೂಚನೆಗಳು

1. ಪ್ರತಿಯೊಂದು ಪ್ರಶ್ನೆಗೆ ಒಂದೇ ಒಂದು ಅತ್ಯಂತ ಸೂಕ್ತವಾದ/ಸರಿಯಾದ ಉತ್ತರವಿರುತ್ತದೆ.
2. ಪ್ರತಿ ಪ್ರಶ್ನೆಗೆ ಒಂದು ವೃತ್ತವನ್ನು ಮಾತ್ರ ನೀಲಿ ಅಥವಾ ಕಪ್ಪು ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್‌ನ್ನಿಂದ ಮಾತ್ರ ತುಂಬತಕ್ಕದ್ದು. ಉತ್ತರವನ್ನು ಮಾರ್ಪಡಿಸಲು ಪ್ರಯತ್ನಿಸಬೇಡಿ.
3. ವೃತ್ತದೊಳಗಿರುವ ಅಕ್ಷರವು ಕಾಣದಿರುವಂತೆ ವೃತ್ತವನ್ನು ಸಂಪೂರ್ಣವಾಗಿ ತುಂಬುವುದು.
4. ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿ ಯಾವುದೇ ಅನಾವಶ್ಯಕ ಗುರುತುಗಳನ್ನು ಮಾಡಬೇಡಿ.
5. ಉತ್ತರಿಸಿದ ಪ್ರಶ್ನೆಗಳ ಒಟ್ಟು ಸಂಖ್ಯೆಯನ್ನು O.M.R. ಹಾಳೆಯಲ್ಲಿ ನಿಗದಿಪಡಿಸಿರುವ ಜಾಗದಲ್ಲಿ ನಮೂದಿಸತಕ್ಕದ್ದು, ಇಲ್ಲವಾದಲ್ಲಿ O.M.R. ಹಾಳೆಯನ್ನು ಮೌಲ್ಯಮಾಪನಕ್ಕೆ ಪರಿಗಣಿಸುವುದಿಲ್ಲ.

Note : English version of the instructions is printed on the front cover of this booklet.